For the molecules below:
- Using the correct combination of in plane, wedged and dashed lines re-draw the compounds below in three-dimensional form.
- Check the boxes below the molecule to indicate whether the compound has a molecular dipole moment ($\mu=0$ or $\mu \neq 0$).
- For those compounds that have dipole moments indicate the direction of the dipole.

\[
\begin{align*}
\text{CH}_3\text{Br} & & \text{CH}_3\text{OCH}_3 \\
\square \mu = 0 & \square \mu \neq 0 & \square \mu = 0 & \square \mu \neq 0 \\
\text{BF}_3 & & \text{NH}_3 \\
\square \mu = 0 & \square \mu \neq 0 & \square \mu = 0 & \square \mu \neq 0
\end{align*}
\]

For the molecules below:
- Check the boxes below the molecule to indicate whether the compound has a molecular dipole moment ($\mu=0$ or $\mu \neq 0$).
- For those compounds that have dipole moments indicate the direction of the dipole.

\[
\begin{align*}
\text{Br} & & \text{Cl} \\
\square \mu = 0 & \square \mu \neq 0 & \square \mu = 0 & \square \mu \neq 0
\end{align*}
\]