

CHE 362L

Physical Chemistry Laboratory 2H

SPRING 2004

All items on this syllabus are subject to change by the Instructor

Instructor: W. Vernon Hicks 572-5406 hicks@nku.edu
Office: SC 447
Office Hours: 9:10 - 11:00 MF, 10:00 – 11:00 W, Other times by appointment
Corequisite/Prerequisite: CHE 361
Credit Hours: 2
Classroom: Sc 427 8:00 -12:50 T
Text: *Experiments in Physical Chemistry, 7th ed.*, by C. W. Garland, J. W. Nibler, and D. P. Shoemaker, McGraw-Hill Companies, Inc., New York, New York, 2003.
Notebook: Lab Record Book for Northern Kentucky University Department of Chemistry, Hayden-McNeil Publishing. One required for each laboratory group.
References: *Statistical Treatment of Experimental Data*, by Young
Data Analysis: Graphing and Report Writing, by Hawkins & Niewahner
Data Analysis Software: *Graphical Analysis, Mathematica, KaleidaGraph*
The student is expected to read the pamphlet *Safety in Academic Chemistry Laboratories* and abide by the safety rules and guidelines contained therein. An abbreviated list of NKU Safety Rules will be provided. Experiments may involve texts and sources other than the required text for the course. Students will work in groups, with typically two students per group.

Tentative list of Experiments:

1. Heat of Combustion.
2. Heat of Neutralization.
3. Two-Component Liquid-Solid Phase Diagram.
4. Effect of Ionic Strength on Solubility.
5. Three-Component Solubility Diagram.
6. Reaction Kinetics. Inversion of Sucrose.
7. Solubility Products from EMF Measurements.
8. Electrical Conductance.
9. IR Spectrum of HCl or HBr.
10. Molecular Calculations.
11. Reaction Kinetics: The Effect of Temperature on Reaction Rate.

The student must do 8 of the above experiments, including 1, 2, 6, 9, and 10. The experiments need not be performed in the order given. The instructor reserves the right to substitute new experiments for some of those given above.

The student will be graded on laboratory technique, the laboratory notebook, and the laboratory reports. Technique will be judged on the basis of direct observation as well as precision and accuracy of the reported results. Each member of a laboratory group is expected to contribute substantially to the performance of the assigned experiments. Some of the criteria that will be considered in grading the laboratory report are data presentation, data manipulation, discussion of results (including analysis of errors), proper literature citation, adherence to proper format, grammar, and conciseness. Plagiarism will not be tolerated. The overall grade for the experiment will be displayed on the returned laboratory report. Points will be deducted for lab reports turned in late. The course grade will be based on the average of the best seven of the eight lab reports. Students may be assessed for broken or damaged equipment.

All raw data must be recorded in the lab notebook for the group and dated.

Lab Report: Each member of a laboratory group must turn in a separate formal laboratory report. The laboratory report must be typed or printed from a word processor. The laboratory report should have the following format:

Cover Page: The cover page must include the following: the title of the experiment, the course title, your name, the names of your lab partners, the beginning and ending date for the experiment, and the date the report is submitted.

Abstract: The abstract must be on a separate page following the cover page. It should be a brief summary of the experiment, including any numerical results and a comparison with literature values, if possible. One or two paragraphs, each with a few sentences, should suffice.

Main Body: The main body must include the objective of the experiment, a brief description of the theory and the experimental procedure in your own words, experimental observations, the raw data collected (preferably in tabular form), and the values calculated from the raw data. Be sure to include units for all data and calculated values, and a precision index for the calculated values.

Sample calculations must be included to show the data manipulation.

A discussion of the results of the experiment is a crucial part of the report. This should include comparisons with literature values (properly referenced with footnotes, *etc*) as well as a discussion of possible errors and their probable effects on the experimental results. Suggestions for improvements can be included.

Conclusions drawn from the results of the experiment must be explicitly stated; these conclusions should refer to the stated objectives of the experiment.

Although it is allowed (even encouraged) for all the members of a group to work up the data together, the error analysis and conclusions should be your own.

References: A list of references must be included on a separate sheet. These should include the source or sources for the experimental procedure followed as well as references for literature values used for comparison with the experimental results obtained.

Lab Report	Due Date
First	Feb 6
Second	Feb 20
Third	March 5
Fourth	March 19
Fifth	March 26
Sixth	April 2
Seventh	April 16
Eighth	April 30

Students are responsible for reading and understanding all items on this syllabus. Any items not understood must be brought to the attention of the Instructor within the first two weeks of class. Students are obligated to follow the Student Honor Code. The Honor Code can be accessed at

<http://www.nku.edu/~deanstudents/HonorCode.htm> .

SAFETY RULES AND REGULATIONS

You are required to notify your instructor at the beginning of the semester if you have a health problem, such as a seizure disorder, diabetes, etc. in order for your instructor to see that you get the necessary medical attention should the need arise.

1. Wear safety glasses in the laboratory at all times. **Protect your eyes!**
2. Wear suitable footwear (no open sandals) in the laboratory at all times.
3. Never smoke, eat, or drink in the laboratory.
4. Perform only those exercises that are authorized.
5. Report all accidents and injuries immediately to your instructor. An **Incident Report** must be filed. Forms are available in room NS 214.
6. Familiarize yourself with the first aid chart and the location of the fire extinguishers, fire blanket, shower and eyewash.
7. Exercise caution in using glass tubing, especially in inserting it into stoppers. Do not use cracked or broken glass equipment.
8. Never heat liquids in graduated cylinders. When you are heating a test tube, never point its mouth towards anyone nor look down into it.
9. Carefully read all labels on reagent bottles before using; an error may have serious consequences. Use only the quantities and concentrations called for.
10. Never pour water into concentrated acid. Always pour the acid slowly into the water while stirring the mixture constantly.
11. Use the fume hood or fume suction device whenever toxic or irritating gases are likely to be evolved. Never directly inhale vapors. Never taste a chemical.
12. Avoid possible poisoning by contamination through the hands. Always wash hands thoroughly after exposure to hazardous chemicals and as a general rule, after each lab period.
13. Use only glassware that is initially clean.
14. A laboratory apron or lab coat to protect your clothing is a wise precaution.
15. If the fire alarm sounds, turn off the gas and leave the building by the nearest exit. If you are the last one out of the lab, close the door behind you. Wait for your instructor's approval before re-entering the building.

The following laboratory regulations must be observed for efficiency and further safety in the laboratory

1. Never throw solid or liquid materials into sinks; **use the correct waste jars**. Your instructor will point these out.
2. Do not insert spatulas or pipets into reagent bottles. Instead, transfer an approximate amount of the reagent desired to a glazed paper, a watch glass, or a test tube. Do not carry stock bottles to your desk.
3. Weigh solids on a glazed paper or on a watch glass; do not allow chemicals to come into contact with the balance pans.
4. Clean up spilled materials immediately, using liberal quantities of water. Neutralize strong acids with sodium bicarbonate, and strong bases with acetic acid, then wash with water.
5. Keep your working surfaces clean at all times. Use a damp sponge for wiping the surface.

6. Before leaving the laboratory, make sure the water and gas are completely shut off. Return all special equipment to the stockroom or designated place.